THE ROLE OF SYNCHRONIZATION OF ATOMIC VIBRATIONS AND ENERGY SELF-REGULATION IN THE FORMATION OF EUTECTIC MATERIALS UNDER CONTACT MELTING CONDITIONS BY

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A hypothetical model of the initial stage of contact eutectic melting is proposed, explaining the phase transition through the synchronization of oscillations of atoms at the phase boundary, with a change in the contact temperature. Synchronization leads to a resonant increase in the amplitudes of atomic (molecule) oscillations, a decrease in their bond energy, and a transition to the liquid phase. The model is consistent with the classical phase diagram, where the eutectic point corresponds to the energetically favorable state of resonant melting. To analyze the eutectic behavior, heterogeneous systems were studied: non-metallic H_2O -NaCl and metallic Ga—In and Al—Cu, in which intermetallic phases are formed. In all cases, a characteristic drop in temperature and transition to a liquid state upon reaching the eutectic are observed. The results confirm that local oscillatory processes play a key role in the formation of a liquid layer of eutectic composition, complementing traditional thermodynamic concepts.

Keywords: eutectic, oscillation synchronization, resonant melting, phonon modes, contact melting

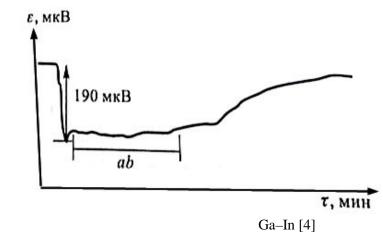
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[1, 2].
             [3, 4].
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                                                        F(T)=U(T)-T\cdot S(T),
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                                    , U-
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(A)[7]
                                                                                                                                              (2)
                                                          <sub>0</sub> - ), F<sub>0</sub> -
       , m –
                                                           0)
                                                            =\frac{1}{2}kA^2 \ge E
                                                                                                                                              (3)
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     k-
                                                                          ),
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[8].
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[6, 9]. G=SL+LV-SV, (4) SL, LV, SV -[10]. G < 0 $\frac{d\theta_i}{dt} = \omega_i + \frac{K}{N} \sum_{j=1}^{N} sin(\theta_i - \theta_i).$ (5) , K – , $_{\rm i}$ $^{\rm -}$ $^{\rm K}$ [11]. H₂O-NaCl. [12]. (~ −21.2 °C). $\mu_{H2\mathrm{O}}$ G= H-T S<0, (6) S>0, [2], Ga-In. 22 °C. In Ga (15.6 °C), 22 °C, 1),

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Cu-Al.

, CuAl₂), ,

550 °C, Cu (1085 °C) [4].

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